Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Class \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Macromolecules – Use Ch 6, Section 4

|  |  |
| --- | --- |
|  CarbohydratesAll macromolecules…Definition:Contain what element?Also known as?Building Blocks (Monomers): Sugars (monosaccharides)Function:Examples:Additional Info: |  LipidsBuilding Blocks:Function:Examples: Additional Info: |
|  ProteinsBuilding Blocks:Function:Examples:Additional Info: |  Nucleic Acids Building Blocks:Function:Examples:Additional Info: |

|  |
| --- |
| Life’s Macromolecules |
| **Organic Compound** | **Use by Organisms** | **Building Blocks** | **Examples** |
| Carbohydrates |  |  |  |
| Lipids |  |  |  |
| Nucleic Acids |  |  |  |
| Proteins |  |  |  |

*Which food molecule (simple sugar carbohydrate, complex sugar carbohydrate, lipid, protein) would you eat if…*

1. …you needed a quick boost of energy?

2. …you wanted to grow healthy hair & nails?

3. …you haven’t eaten in days?

4. …you wanted to grow healthy hair?

5. …you had a race tomorrow afternoon?

6. …you were getting ready for hibernation?

7. …you wanted to get bigger muscles?

Water and pH Notes Organizer

1. What is a ***polar molecule***?
2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ is an important polar molecule for all living things.
3. Water is held together by what *unique and weak interaction*?
4. Illustrate a **water molecule**.
5. Water is an extremely useful molecule because water is both **cohesive** and **adhesive**. Define the following terms and give an example of each:

	1. **Cohesion**
	2. **Adhesion**
6. Explain what it means to describe water as a “***universal solvent***.”
7. What is *evaporative cooling*?
8. Water \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ during freezing.
9. **Homogeneous mixtures** have what type of composition?

	1. ***Solvent***:
	2. ***Solute***:
10. Describe a **heterogenous mixture** and provide an example.
11. What is **pH**?
12. What is the pH of an ***acidic*** solution?
13. What is the pH of a ***basic*** solution?
14. Define the following terms.

	1. **Acids**:
	2. **Bases**:
15. Describe the pH of *water*.